

Key:-

Chapter 17:

Terminology:

Term	definition	keyword
Radioactivity	The emission of tiny indivisible particles by the nuclei of certain atoms	
Isotopes	Atoms w/ different mass # but same atomic #	
Nuclide	The nucleus of a specific atom	
Radioactive decay	Nuclear decay	
Positron emission	inc # of neutrons, dec # of pro	
Electron capture	Nuc. absorbs <del>the</del> elect.	
Half-Life	The period of time needed for 1/2 of a sample of radioactive atoms of a given isotope to decay	

Questions:

1. What holds the nucleus together?

Strong nuclear force

2. What is the force of repulsion between the protons in an atom?

Electrostatic repulsion

3. Describe the 3 radioactive particles from greatest penetration to least penetration.

Chapter 18: On back

What is the difference between saturated and unsaturated hydrocarbons?

Sat - every C bonded

Unsat - Contains double + triple bonds

Describe the following types of hydrocarbons:

Alkanes:  
- Contain single bonds only  
- Saturated

Alkenes:  
- Contain at least 1 double bond  
- Unsat.

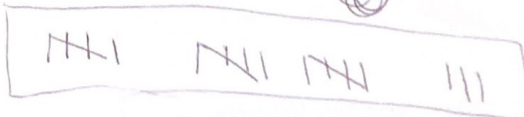
Alkynes: Contain @ least 1 triple bond  
- Unsat

What are isomers?

Compounds w/ the same molecular formula but different structural formulas.

What are Normal Hydrocarbons?

Composed of straight chains of C atoms w/o branching.



3. gamma rays - greatest pen  
lowest ion.

Beta Particles - intermediate ionizing  
- intermed. Pen

Alpha Particles  
- greatest ionizing  
- lowest pen.